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 -- --x = nc V ×25.0
 mL OH n - =12.5
 mmol HCl(aq) HCl(aq)
 [HCl(aq)] 12.5 mmol =
 = n V 0.10 mmol
 HCl(aq) /mL V =120 mL
 2. (a) Given: V HCl(aq)
 = 25.00 mL; [HCl(aq)]
 = 0.350 mol/L; V
 NaOH(aq) = 5.0 mL;
 [NaOH(aq)] = 0.500
 mol/L; Required:
 [H+(aq)], pH Analysis:
 Before titration,

amount of H⁺ is: H(aq)
 H(aq) H(aq) 0.350
 mmol mL ++ +=x =
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 $Q = 0.020$ Statement:
 The value of Q is 0.020. Q is equal to K , so the reaction is at equilibrium. (b) Given: $[HI(g)] = 0.20 \text{ mol/L}$; $[H_2(g)] = 0.15 \text{ mol/L}$; $[I_2(g)] = 0.090 \text{ mol/L}$; $K = 0.020$ Required: Q ; direction of reaction to reach equilibrium
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Solution: OH OH OH
0.50 mmol mL --
-- = x = nc V × 25.0 mL
OH n - = 12.5 mmol
HCl(aq) HCl(aq)
[HCl(aq)] 12.5 mmol =
= n V 0.10 mmol
HCl(aq) /mL V = 120 mL
2. (a) Given: V HCl(aq)
= 25.00 mL; [HCl(aq)]
= 0.350 mol/L; V
NaOH(aq) = 5.0 mL;
[NaOH(aq)] = 0.500
mol/L; Required:
[H+(aq)], pH Analysis:

Before titration,
amount of H+ is: H(aq)
H(aq) H(aq) 0.350
mmol mL ++ += x =
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 $(0.040)(0.010) / (0.14)^2$

$Q = 0.020$ Statement:

The value of Q is
 0.020 . Q is equal to K ,
so the reaction is at
equilibrium. (b) Given:

$[HI(g)] = 0.20 \text{ mol/L}$; $[H_2(g)] = 0.15 \text{ mol/L}$; $[I_2(g)] = 0.090 \text{ mol/L}$; $K = 0.020$ Required: Q ;

direction of reaction to
reach equilibrium

Solution: $Q = \frac{[H_2(g)][I_2(g)]}{[HI(g)]^2} =$
 $(0.15)(0.090) / (0.20)^2$
 $Q = 0.34$

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