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two H's (hydrogen atoms) in each molecule of H₂O. 2. Coefficients - the regular-sized numbers that tell you how many molecules you have. For example, in 3H₂O, the 3 is the coefficient. This shows that there are three H₂O molecules in

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 History of the Atom
 Democritus- Greek philosopher; 400 BC 1st to use word "atomos" (means ____) Aristotle-refuted what Democritus said; all matter comes from "hyle" Early Alchemists-1200-1500's; tried to turn material into ____; kept chemistry alive;

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 Guided Notes: Atomic Structure 01 Atom: the smallest particle of an element that retains its ____ in a ____ . Democritus believed that atoms are ____ & ____ Think about it: Was Democritus correct? Two issues with Democritus's ideas: 1. His ideas did not explain ____.

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Notes on Thomson's Model of the Atom
 Original Theories of Matter: 1) Democritus: matter is made up of indivisible particles called atoms 2) Dalton: elements, compounds, law of conservation of mass Which of the above theories of matter needs to be revised to become more accurate now that we know the results of Thomson's experiments with electrons?

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 Glossary TERM DEFINITION anion an ion that has a negative charge atom atomic mass atomic weight the average atomic mass of a naturally occurring sample of an element cation compound a substance made up of two or more chemical elements that are chemically bonded together electron element energy level ion isotope mass number the total number of protons and neutrons in an atom neutron a subatomic particle that has about the same mass as a proton but does not have an

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Guided Notes: Atomic Structure 01
Atom: the smallest particle of an element that retains its _____ in a _____.
Democritus believed that atoms are _____ & _____. Think about it: Was Democritus correct? Two issues with Democritus's ideas: 1. His ideas did not explain _____.

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The Structure of the Atom Glossary
TERM DEFINITION anion an ion that has a negative charge atom atomic mass atomic weight the average atomic mass

of a naturally occurring sample of an element cation compound a substance made up of two or more chemical elements that are chemically bonded together electron element energy level ion isotope mass number the total number of protons and neutrons in an atom neutron a subatomic particle that has about the same mass as a proton but does not have an ...

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Guided Notes on Thomson's Model of the Atom Original Theories of Matter: 1) Democritus: matter is made up of indivisible particles called atoms 2) Dalton: elements, compounds, law of conservation of mass Which of the above theories of matter needs to be revised to become more accurate now that we know the results of Thomson's experiments with electrons?

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